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For this chapter, you will need to know about kinetic energy. Kinetic energy is what is created when an object moves. Kinetic theory also states that all matter consists of particles that are constantly in motion. Therefore, when in a closed container, all these moving particles

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that make up, well,
everything will have to
hit each other.

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The Nature of Gases

OBJECTIVES: Describe
the assumptions of the
“kinetic theory” as it
applies to gases. 3.

Section 13.1 The

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OBJECTIVES: Study
Interpret gas pressure
in terms of kinetic
theory. 4.

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particles have
attraction forces
between them so they
can't spread far away

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Chapter 12: states of
matter Gases expand,
diffuse, exert pressure,

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Chapter 13 States

and can be compressed because they are in a low-density state consisting of tiny constantly-moving particles. Intermolecular forces, including dispersion forces, dipole-dipole forces, and hydrogen bonds, determine a substance's state at a given temperature.

**Chapters 12 & 13:
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